

**B TECH**  
**(SEM I) THEORY EXAMINATION 2018-19**  
**CHEMISTRY**

Time: 3 Hours

Total Marks: 100

Note: 1. Attempt all Sections. If require any missing data; then choose suitably.

**SECTION A**

1. Attempt all questions in brief. 2 x 10 = 20
- Which of the following molecules will show rotational spectrum:  $H_2$ ,  $HCl$ ,  $CH_4$ ,  $CH_3Cl$ ,  $CH_2Cl_2$ ,  $H_2O$  and  $SF_6$ ?
  - What is meant by the term polarizability in Raman spectra?
  - Explain why the value of GCV is greater than NCV.
  - Why does magnesium bicarbonate require double amount of lime for softening?
  - Write the monomer of a) Neoprene b) Terylene.
  - Why dry ether solvent important for the preparation of Grignard reagent?
  - Comment on the use of aluminum in place of Zinc for Cathodic protection of iron from rusting.
  - Calculate the cell potential at 298 K for the reaction:  
 $Al^{3+} + Fe \rightarrow Al + Fe^{3+}$   $E^0_{cell} = -1.62$   
The concentration of  $Al^{3+}$  and  $Fe^{3+}$  are 1.2 and 2.5 M.
  - Write any two applications of Nanotechnology.
  - Arrange the following molecules/ions in order of their increasing bond length:  $N_2$ ,  $N_2^+$ ,  $N_2^{2-}$ .

**SECTION B**

2. Attempt any three of the following: 10 x 3 = 30
- Discuss the structure, preparation, properties and applications of fullerenes?
  - Derive Nernst Equation. The voltage of the cell  $Pb/PbSO_4/Na_2SO_4/Hg$  is 0.9647 V at 25°C the temperature is  $1.74 \times 10^{-4} \text{ VK}^{-1}$ . Calculate the values of  $\Delta G$ ,  $\Delta S$  and  $\Delta H$ .
  - What is finger print region and functional group region in IR spectroscopy? Two isomers A and B of the molecular formula  $C_3H_6O$  gives IR absorption at  $1650 \text{ cm}^{-1}$  and  $1710 \text{ cm}^{-1}$  respectively. Assign structural formula to A and B isomers?
  - What is hardness of water? What do you mean by term permutit? Describe Zeolite or Permutit process for softening of hard water.
  - What are organometallic compounds? Explain various methods of preparation of Grignard reagent and also write reactions of Grignard reagent with  $HCHO$ ,  $R_2NH$ ,  $CO_2$ ,  $CH_3CH_2OH$  and ester.

**SECTION C**

3. Attempt any one part of the following: 10 x 1 = 10
- What are liquid crystals? Distinguish between Nematic & Smectic liquid crystal and write the applications of liquid crystal?

- (b) Explain BMO and ABMO and differentiate between them. Draw molecular orbital diagram of  $\text{NO}^+$ . Calculate its bond order and predict its magnetic properties.

**4. Attempt any one part of the following: 10 x 1 = 10**

- (a) State the selection rule for Raman spectroscopy. What technological advances have enabled the routine use of Raman spectroscopy? Which of the following spectroscopy (IR or Raman) would you use to measure the vibrational frequency of the following bonds:

- i) The stretching frequency of  $^{14}\text{N}-^{15}\text{N}$   
ii) The  $\text{C}\equiv\text{C}$  str in Ethyne,  $(\text{CH}\equiv\text{CH})$   
iii) The  $\text{C}=\text{O}$  str in acetone,  $\text{CH}_3\text{COCH}_3$   
iv) The  $\text{Re}-\text{Re}$  str in compound,  $(\text{CO})_5\text{Re}-\text{Re}(\text{CO})_5$

- (b) What is Beer-Lambert law in UV-Visible absorption spectroscopy? A compound having concentration  $10^{-3}\text{g/l}$  resulted absorbance value 0.20 at  $\lambda_{\text{max}}$  510 nm using 1.0 cm cell. Calculate its absorptivity and molar absorptivity values. Molecular weight of compound is 400. Can ultra-violet spectral data be useful to distinguish the following compounds? Give reasons.

- (i) Ethyl benzene and styrene.  
(ii)  $\text{CH}_2=\text{CH}-\text{CH}_2-\text{CH}=\text{CH}_2$  and  $\text{CH}_2=\text{CH}-\text{CH}=\text{CH}-\text{CH}_3$

**5. Attempt any one part of the following: 10 x 1 = 10**

- (a) Define and explain the terms involved in phase rule. Draw a neat labeled phase diagram of water system and calculate degree of freedom of areas and curves in it. What is the significance of the triple point and metastable curve in the system?
- (b) Describe the mechanism of electrochemical or wet corrosion with help of reactions? Explain the cathodic protection method of prevention of corrosion.

**6. Attempt any one part of the following: 10 x 1 = 10**

- (a) Explain reverse osmosis with its advantages. A water sample on analysis gives the following data:  $\text{Ca}^{2+}=20$  ppm,  $\text{Mg}^{2+}=25$  ppm,  $\text{CO}_2=30$  ppm,  $\text{HCO}_3^- =150$  ppm,  $\text{K}^+ =10$  ppm. Calculate the lime (87% pure) and soda (91% pure) required to soften  $10^6$  liter of sample water.
- (b) Describe proximate analysis of fuels. A coal sample has the following composition by weight:  $\text{C}=90\%$ ,  $\text{O}=3\%$ ,  $\text{S}=0.5\%$ ,  $\text{N}=0.5\%$  and  $\text{Ash}=2.5\%$ . Net calorific value of the coal was found to be 8,490.5 kcal/kg. Calculate the percentage of hydrogen and Gross calorific value.

**7. Attempt any one part of the following: 10 x 1 = 10**

- (a) Write a note on (i) Polymer blends (ii) Polymer composites
- (b) Differentiate between elastomers and fibers? Give the preparation, properties and uses of Buna-S, Buna-N and Neoprene.